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### **Section 7 1 Describing Reactions**

192 Chapter 7 7.1 Describing Reactions  
Reading Strategy Monitoring Your  
Understanding Preview the Key  
Concepts, topic headings, vocabulary,  
and figures in this section. List two  
things you expect to learn. After

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reading, state what you learned about each item you listed. Key Concepts What is the law of conservation of mass? Why must chemical

### **Section 7.1 7.1 Describing Reactions - PC\|MAC**

Section 7.1 Describing Reactions. (pages 192-198) This section discusses the use

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of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry. Reading Strategy (page 192) Monitoring Your Understanding Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn.

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### **Chapter 7 Chemical Reactions**

#### **Section 7.1 Describing Reactions**

Section 7.1 Describing Reactions (pages 192–198) This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry. Reading Strategy (page 192) Monitoring Your Understanding Preview the Key



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Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn.

### **Chapter 7 Chemical Reactions**

#### **Section 7.1 Describing Reactions**

Use Figure 2 to summarize the key concepts in this section, such as identifying reactants and products,

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balancing equations, understanding conservation of mass, and determining mole ratios. Have students determine the mole ratio of carbon to carbon dioxide in the balanced equation. (1:1) Solutions.

### **Section 7.1 7.1 Describing Reactions - Physical Science**

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7.1 Describing Reactions 10 Terms.  
phoebexhuang. Glencoe Physical  
Science Chapter 19 (Chemical  
Reactions) 9 Terms. Roman\_Slaves.

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Pearson Physical Science Concepts in Action Chapter 7 Section 1 14 Terms. graciejones12. OTHER SETS BY THIS CREATOR. Word of the day Final Study Guide 28 Terms. wdesrochers. Freshman English Rhetorical Appeals & Irony ...

### **Chapter 7.1 Describing Reactions Flashcards | Quizlet**

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The substances undergoing reaction are called reactants, and their formulas are placed on the left side of the equation. The substances generated by the reaction are called products, and their formulas are placed on the right side of the equation.

### **7.1: Writing and Balancing Chemical**

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### **Equations | General ...**

physical science chapter 7: describing reactions. Reactants. Products. Chemical equation. Carbon and oxygen react to form carbon.... The substances that undergo change. The new substances formed as a result of that change.

Representation of a chemical reaction in which the reactants a.... Say  $C+O_2$ ---

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CO<sub>2</sub> in words.

### **describing reactions chapter 7 Flashcards and Study Sets ...**

Section 7.1 - Describing Reactions.  
Chemical Equations. A useful description of a chemical reaction tells you the substances present before and after the reaction. In a chemical reaction, the



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## Describing Reactions Answers

substances that undergo change are called . reactants. The new substances formed as a result of that change are called . products.

### **Chapter 7 Notes - Chemical Reactions**

Chapter 7 Notes - Chemical Reactions  
Section 7.1 - Describing Reactions

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## Describing Reactions Answers

Chemical Equations A useful description of a chemical reaction tells you the substances present before and after the reaction.-In a chemical reaction, the substances that undergo change are called reactants.

**Chapter 7 Notes- General Physical Science.doc - Chapter 7 ...**

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inc documentary discussion Page 1/3.

## **Section 7 1 Describing Reactions Answers**

A solution is prepared by dissolving 44.6 grams of acetone ( $\text{OC}(\text{CH}_3)_2$ ) in water to produce 1.50 Liters of solution. What is the molarity of the resulting solution? A certain laboratory procedure requires

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0.025 M  $\text{H}_2\text{SO}_4$ . How many milliliters of 1.10 M  $\text{H}_2\text{SO}_4$  should be diluted in water to prepare 0.500 L of 0.025 M  $\text{H}_2\text{SO}_4$ ?

### **4.9: Describing Reactions in Solution - Chemistry LibreTexts**

A. Describe each reaction. Explain what changes were signs of a chemical

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reaction. B. Which reaction was endothermic, and which was exothermic? How could you tell?

Conclusion A. Paragraph 1: purpose, materials, procedures (2 points) B. Paragraph 2: data. Describe the reactions, and explain which reaction is exothermic and which is endothermic ...

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### **Chapter 7: Chemical Reactions**

reactions are synthesis reactions, decomposition reactions, single-replacement reactions, double-replacement reactions, and combustion reactions. Each type describes a different way in which reactants interact to form products. Chemical Reactions 199  
FOCUS Objectives 7.2.1 Classify



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## Describing Reactions Answers

chemical reactions as synthesis, decomposition, single-replacement, double-

### **7.2 Types of Reactions - Physical Science**

Chapter 7: Chemical Reactions

DISCUSSION PLAN Chapter Summary

Chapter 7 reviews the information about

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## Describing Reactions Answers

chemical changes and explains how a chemical reaction can be written down as a chemical equation. The concept of strong electrolytes is introduced to explain solubility and precipitation reactions. Precipitation reaction is used to demonstrate types of chemical equations: molecular, complete ionic, and net ionic equations.

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### **Chapter 7 Answer Key - Chapter 7 Chemical Reactions ...**

3/24/2015 2. Chemical Reactions. • It is a change in matter that produces one or more NEW SUBSTANCES with different properties. • A chemical reaction is another name for a chemical change • New properties are formed • No new

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atoms are made, but atoms are rearranged • New compounds can be made • Old bonds are broken and new bonds are formed • Always involves changes in energy.

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