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= 12.00 mmol - 5.40 mmol Unreacted
ethanoic acid = 6.60 mmol Since 18.00
mL of NaOH(aq) was added to 20.00 mL
of ethanoic acid solution, the total
volume is now 38.00 mL. Solution: The
concentration of unreacted ethanoic
acid, [CH₃COOH(aq)], can be
determined: $\frac{6.60 \text{ mmol}}{38.00 \text{ mL}} = 0.174 \text{ mol/L}$

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$\text{HC}_2\text{H}_3\text{O}_2(\text{aq})$ 23.2 [$\text{HC}_2\text{H}_3\text{O}_2(\text{aq})$] (6.60
mmol) = 38.00 mL

Section 8.7: Acid-Base Titration Tutorial 1 Practice, page 547

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