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1/T (K<sup>-1</sup>) 0.00180 0.00174 0.00159  
0.00150 0.00143 0.00128 ln k -10.02  
-8.92 -5.98 -4.26 -2.74 0.29 Gradient =  
-19897 = -E<sub>a</sub> / R E<sub>a</sub> = 19897 × 8.31 =  
165000 J mol<sup>-1</sup> = 165 kJ mol<sup>-1</sup> 6  
Temperature (°C) 20 50 100 150 200  
250 k (mol<sup>-1</sup> dm<sup>3</sup> s<sup>-1</sup>) 0.0147 0.0675  
0.497 2.29 7.61 20.1 1/T (K<sup>-1</sup>) ln k  
Gradient = -4812 = -E<sub>a</sub> / R

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TASK 1 - Bronsted-Lowry acids & bases

1 Acid =  $\text{H}_2\text{O}$ , base =  $\text{NH}_3$  2 Acid =

$\text{HCl}$ , base =  $\text{H}_2\text{O}$  3 Acid =  $\text{HCOOH}$ , base

=  $\text{KOH}$  Chemsheets A2 1156 (QC - Kp A)

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© wwwCHEMSHEETS.co.uk 3-Mar-2018  
Chemsheets A2 1156 Nitrogen reacts  
with hydrogen as shown:  $3\text{H}_2(\text{g}) + \text{N}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$   $\Delta H = -76 \text{ kJ mol}^{-1}$  100  
moles of

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## **A level Chemistry - Kinetics |**

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## Teaching Resources

- 1) Increase the temperature Number of particles with  $E \geq E_a$  more same less
- 2) Remove half the molecules Number of particles with  $E \geq E_a$  more same less Energy
- 4) Reduce the volume of the container Number of particles with  $E \geq E_a$
- 3) Add a catalyst Number of Energy particles with  $E \geq E_a$  more same less

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Energy 6) Add an inert gas more am less  
Energy

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1) What is ligand substitution • A ligand substitution reaction is one where one ligand is replaced by another ligand. 2) Substitution by similar sized ligands

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1)  $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$   
initial moles 1 1 0 0 change in moles  
equilibrium moles 0.4 9) 2)  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$   
initial moles 0.1 0.1 0

## **EQUILIBRIUM QUANTITIES 1**

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## **Worksheet 14: Chemical Kinetics - Chemistry LibreTexts**

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Chemsheets A2 029 9 IR TASK 2 1) The  
IR spectra of six compounds are shown.  
The compounds are: butanoic acid  
butanone but-3-en-1-ol

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## **Chemsheets A2 029 (Spectroscopy) - Weebly**

1) Propene reacts with HBr to form H. H reacts with sodium hydroxide to form I, and I reacts with warm acidified potassium dichromate (VI) to form J. The infra-red spectra of H, I and J are given below, but it does indicate which is - which.

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## **IR TASK 1 - Weebly**

A rate is a measure of how some property varies with time. Speed is a familiar rate that expresses the distance traveled by an object in a given amount of time. Wage is a rate that represents the amount of money earned by a person working for a given amount of

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time.

## **17.1: Chemical Reaction Rates - Chemistry LibreTexts**

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- 2 Acid =  $\text{HCl}$ , base =  $\text{H}_2\text{O}$
- 3 Acid =  $\text{HCOOH}$ , base =  $\text{KOH}$
- 4 Acid =  $\text{HCl}$ , base =  $\text{CH}_3\text{COOH}$
- 5 Acid =  $\text{HCl}$ , base =  $\text{NH}_3$
- 6 Acid =  $\text{HCO}_3^-$ , base =  $\text{OH}^-$
- 7 Acid =  $\text{H}^+$ , base ...

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## **Organic Nomenclature 3 Answers Chemsheets**

Example 1 Imagine a reaction where R reacts with S to make some products:  $R + S \rightarrow \text{products}$  If we compare experiments 1 and 2: Initial [R]  
Experiment mol dm<sup>-3</sup> 1.0 2.0 Initial [S]  
mol dm<sup>-3</sup> 9.0 1.0 '2.0 The concentration of R has been doubled, but the

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concentration of S remains the same. the rate has doubled.

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